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Chapter 2 : Chemical element - Wikipedia

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Chemistry in its element: Know what it is yet? Peter Wothers In , Madame Curie hired a newly-qualified, twenty-year old technician, Marguerite Catherine Perey, to act as her lab assistant. Ten years later, this remarkably skilled woman discovered the much sought after element francium. As a result, she was encouraged to study for a degree, and then for her PhD. Despite the fact that her mother was convinced she would fail, in March Perey successfully defended her thesis on Element Sixteen years later, she became the first woman to be elected a member of the French Academy of Sciences, an honour not even awarded to her mentor Madame Curie. But the story of element 87 begins much earlier than its discovery date in When Mendeleev first proposed his periodic table in he left gaps for elements not yet discovered, but that he predicted should exist. One such gap was one beneath caesium, a position later found to belong to francium. More than 40 years later, it was suggested that each element has a unique numbered position in the periodic table, known as its atomic number. Physicist Henry Moseley proved the existence of the "atomic number" and also suggested that it represented the number of positively-charged protons in the nucleus of the atom. Once the atomic numbers for all the known elements were assigned, it was clear seven elements were missing from the periodic table between hydrogen with atomic number 1, and uranium, number Francium, number 87, was the last of these elements to be discovered in nature. From its position in the table, it was clear that element 87 would be a reactive alkali metal, the heaviest member of the family lithium, sodium, potassium, rubidium and caesium. Consequently, many researchers started looking for the new metal in ores which contained these related elements. Many false claims were made before it was realised that the missing element would be radioactive with no stable form. Then the search focused on looking at the decay sequences of other radioactive elements. Two simple rules dictate what elements are formed in a decay series. For each alpha particle a radioactive sample emits, the atomic number of the product element formed is two less than the element from which it formed. For each beta particle emitted, the atomic number of the product increases by one. The element with atomic number 89, two places to the right of our 87, is actinium which had been discovered in This then decays by emitting an alpha particle to form element 88, radium. This was made even more difficult to spot by the fact francium has a half life of just 21 minutes, because it quickly emits a beta particle to once again form radium. During these initial investigations, Perey referred to her element as actinium-K; a reference to the route by which it is formed. However, she needed a proper name for her element. During her PhD exam, she suggested the name "catium" since she thought it would be the metal that most readily loses an electron to form a cation. Fortunately, this name was met with little enthusiasm - one of her examiners even suggested that English-speaking people might think it was named after a cat. Perey then suggested the name Francium, after her native country, and this name was accepted. Whilst it is naturally occurring, or to be more precise, naturally formed - albeit briefly - during radioactive decay of other elements, the amount of francium on earth is tiny. On descending a group in the periodic table, on average the outermost electrons get further and further away from the nucleus and as a result, become easier to remove from the atom. This is the trend for the elements lithium, sodium, potassium, rubidium and caesium. However, for the really heavy elements, the presence of so many positively charged protons in the nucleus has the affect of causing the electrons to move round at incredibly fast speeds approaching the sound of light. As Einstein realised at such speeds strange things being to happen. The electrons become a little closer to the nucleus than expected and they also become slightly harder to remove than expected. Remarkably considering its short half-life, it has been possible to measure experimentally the energy needed to remove an electron from francium to form a positively charged francium cation. The energy needed is kJ mol -1 some 17 kJ mol-1 more than for caesium. End promo Help text not available for this section currently Video.

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Chapter 3 : FR: conditional sentence elements in French | WordReference Forums

Combining FormsA combining form is created when a word root is combined with a vowel. The word root plus this combining vowel is what creates the combining form. The combining form is the basic foundation to which other word elements are added to build a complete medical word. p. 3 (table)The.

I have, however, seen bread turned to stone. Years ago when my sister and I were in college, our college class at church had a special turkey dinner for the class. This was no disaster, and no one besides the preparers of the meal ever knew about it. My sister fixed one of my favorite recipes, a recipe which my mother and my wife have successfully used for many years. I can still remember the puzzled looks on the faces of those kids. Some foolishly tried to break them in half. Eventually everyone gave up and left them alone. I actually tried to save one of those petrified rolls for posterity. Turning bread into stone was no miracle, only a mishap which gave us an occasion for a good laugh. In our text, our Lord Jesus Christ is challenged by Satan to turn stone into bread, a miracle indeed. These are not the only temptations which occurred during that 40 day testing period, but they are the three which both Matthew 4: We must therefore conclude that these temptations are of significance to these writers, and thus to the gospel, and ultimately to us. The Importance of the Temptation Account There are several reasons why the temptation accounts are of importance to us. Let us consider these as we seek to prepare our hearts and minds for the instruction God has for us from our passage. Furthermore, if James informs us that God cannot be tempted James 1: To know the mindset and the methods of our enemy, the Devil, we are forewarned and forearmed as to the temptations by which he will seek to destroy us. In the three challenges and solicitations of Satan in the temptation accounts of Matthew and Luke we find the three primary avenues by which Satan seeks to make inroads into our lives so as to devastate our spiritual walk with God through Christ. Finally, be strong in the Lord, and in the strength of His might. Put on the full armor of God, that you may be able to stand firm against the schemes of the devil. For our struggle is not against flesh and blood, but against the rulers, against the powers, against the world forces of this darkness, against the spiritual forces of wickedness in the heavenly places. Therefore, take up the full armor of God, that you may be able to resist in the evil day, and having done everything, to stand firm Eph. It was also the temptation of our Lord as God. The temptations of our Lord were those which could be pressed on one who was divine. It was evident that Satan was the source of the temptation. Our temptations are more indirect, coming most often through the world and the flesh. There was no inner inclination to rebel against God and no inner desire to sin. For us it is entirely a different matter, as Romans chapter 7 makes abundantly clear. Temptation is, on the one hand, a solicitation to sin, to do that which is contrary to the will and the word of God. Temptation is an attempt to cause a person to sin. Thus, in the case of Job cf. These two meanings of the same term have long been recognized by biblical scholars. Both the deity and the humanity of the Lord Jesus have been documented. Jesus was prophesied to be born of a woman, but also a product of the miraculous intervention of the Holy Spirit 1: The account of His birth in chapter 2 shows this to have happened. In chapter 3, John the Baptist began his public ministry, preparing the people for the coming of Messiah, who was greater than he. Besides the testimony of John, the Father and the Spirit bore witness to the identity of the Lord Jesus. The descent of the Holy Spirit was the enduement of power for this task. Thus our Lord is both God and man. As man Jesus was both a descendent of David, but also a son of Adam. Two Assumptions Which Need to be Challenged There are two assumptions which are widely held by Christians which need to be challenged, and at least re-thought. The agony which I do find in the Bible is that of our Lord in the Garden of Gethsemene, when He struggled with the reality of the wrath of God which He was about to experience. Satan is, in psychological terminology, a pathological liar. Such persons lie whether or not it appears necessary, and even when it may prove detrimental. I am not at all certain that just because Satan claims to possess all the kingdoms of the world Luke 4: In my opinion, Satan is always offering others that which he does not possess. For example, he encouraged Adam and Eve to help themselves to the forbidden fruit of the tree of the knowledge of good and evil and thus to a new level of knowledge. Our Lord offers men what He possesses, what He has purchased e. I therefore urge you to be careful about believing anything Satan might say, even in

the presence of God. Our Approach Our approach to the temptation of our Lord will be to study it in several segments. In this lesson we will focus on the setting of the temptation Luke 4: We will then seek to see how the principles which guided our Lord in His response can be found repeatedly in His teaching and ministry. Finally, we will seek to discover the forms in which this same temptation can be identified in our own culture, and how they should be dealt with. It was also in the wilderness that John grew up and ministered 1: So, too, it was in the wilderness that Israel tempted God Ps. The Lord went out to the wilderness to confront Satan, or at least to be confronted by him, and to return victorious. Even in His temptation our Lord is in control, not Satan. Our Lord fasted for forty days and nights Matt. Israel was in the wilderness 40 years, even as our Lord was in the wilderness for 40 days. Israel hungered even as our Lord did. In both cases God was testing man. In the case of Israel, they also put God to the test, demanding to be fed, and sometimes threatening to return to Egypt. Our Lord is the antitype of Israel cp. There were, by inference cf. The scene is therefore set. Make Stone into Bread 4: Our Lord was understandably hungry. More than just hunger is involved, however. We can miss a mere meal and feel a strong sense of hunger. If He continued not eating, He would die. Thus, Satan challenged, He must act. Whether or not He must act as Satan had challenged, by miraculously converting stone to bread, is doubtful, for it would seem that there would have been other means of satisfying His need for nutrition. Satan seems to be appealing to that basic human instinct of self-preservation. Bread was a basic essential of life, not a luxury food item. It is not self-indulgence, then, but self-preservation which Satan is seeking to induce our Lord to accomplish through His divine power. On the face of the matter, such an offer seems innocuous. After all, is there anything so wrong with meeting basic human needs? Hunger is a need which our Lord would later meet in His public ministry. Did He not feed the 5,, who were in the wilderness and without food cf. For our Lord, serious hunger even justified setting aside normal rules. The answer was to be found in the Word of God itself. Jesus responded to Satan in the words of Deuteronomy: Israel was about to enter into the promised land and God was, through Moses, reminding His people of the basis on which His blessings would be bestowed in the land. We have already noted the parallel which the gospel accounts draw between the experience of Israel in the wilderness and that of our Lord in the wilderness. In Deuteronomy God is referring to the lessons which God has taught Israel in order to prepare her for the blessings of the land. If Satan was subtly suggesting that hunger was inconsistent with divine presence and blessing, Jesus knew from Deuteronomy that it was the evidence of divine love and care, in preparation for blessing. The only reason that Jesus did not make the stone into bread was because it would have been wrong to do so. Jesus had the power to change His circumstances, to satisfy His hunger, but He refused to employ it in such a fashion. God made it abundantly clear to Israel that when they complained about their circumstances, they ultimately complained against God, for it was He who led them. For our Lord to have acted as Satan proposed would have been an act of disobedience. Fifth, the only motive for making the stone into bread would have been distrust regarding the goodness and the guidance of God. As I understand the Bible, unbelief is the ultimate root of most, if not all, disobedience. Israel grumbled against God in the wilderness and demanded that God prove Himself because they doubted His goodness and guidance. Sixth, Life is more than mere physical survival and thus must be sustained by more than food. Surely the Old Testament not to mention the New Testament makes this abundantly clear. God told Adam and Eve that they would die if they ate the forbidden fruit, yet they continued to live physically after their disobedience. We know that the death they experienced included physical death, but involved much more. So, too, life was much more than physical existence. Intimacy with God was one of the things which was lost, for the evening walks in the garden were ended, along with life in the garden. So it was for Adam and Eve as well.

Chapter 4 : Francium | Define Francium at calendrierdelascience.com

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Almost all other elements found in nature were made by various natural methods of nucleosynthesis. New atoms are also naturally produced on Earth as radiogenic daughter isotopes of ongoing radioactive decay processes such as alpha decay , beta decay , spontaneous fission , cluster decay , and other rarer modes of decay. Of the 94 naturally occurring elements, those with atomic numbers 1 through 82 each have at least one stable isotope except for technetium , element 43 and promethium , element 61, which have no stable isotopes. Isotopes considered stable are those for which no radioactive decay has yet been observed. Elements with atomic numbers 83 through 94 are unstable to the point that radioactive decay of all isotopes can be detected. Some of these elements, notably bismuth atomic number 83 , thorium atomic number 90 , and uranium atomic number 92 , have one or more isotopes with half-lives long enough to survive as remnants of the explosive stellar nucleosynthesis that produced the heavy metals before the formation of our Solar System. As of , there are known elements in this context, "known" means observed well enough, even from just a few decay products, to have been differentiated from other elements. Six of these occur in extreme trace quantities: These 94 elements have been detected in the universe at large, in the spectra of stars and also supernovae, where short-lived radioactive elements are newly being made. The first 94 elements have been detected directly on Earth as primordial nuclides present from the formation of the solar system, or as naturally occurring fission or transmutation products of uranium and thorium. The remaining 24 heavier elements, not found today either on Earth or in astronomical spectra, have been produced artificially: Technetium was the first purportedly non-naturally occurring element synthesized, in , although trace amounts of technetium have since been found in nature and also the element may have been discovered naturally in The nuclides of stable and radioactive elements are also available as a list of nuclides , sorted by length of half-life for those that are unstable. One of the most convenient, and certainly the most traditional presentation of the elements, is in the form of the periodic table , which groups together elements with similar chemical properties and usually also similar electronic structures. Atomic number Main article: Thus, all carbon isotopes have nearly identical chemical properties because they all have six protons and six electrons, even though carbon atoms may, for example, have 6 or 8 neutrons. That is why the atomic number, rather than mass number or atomic weight , is considered the identifying characteristic of a chemical element. The symbol for atomic number is Z. Isotope , Stable isotope ratio , and List of nuclides Isotopes are atoms of the same element that is, with the same number of protons in their atomic nucleus , but having different numbers of neutrons. Thus, for example, there are three main isotopes of carbon. All carbon atoms have 6 protons in the nucleus, but they can have either 6, 7, or 8 neutrons. Since the mass numbers of these are 12, 13 and 14 respectively, the three isotopes of carbon are known as carbon , carbon , and carbon , often abbreviated to 12C, 13C, and 14C. Carbon in everyday life and in chemistry is a mixture of 12C about Most 66 of 94 naturally occurring elements have more than one stable isotope. Except for the isotopes of hydrogen which differ greatly from each other in relative massâ€”enough to cause chemical effects , the isotopes of a given element are chemically nearly indistinguishable. All of the elements have some isotopes that are radioactive radioisotopes , although not all of these radioisotopes occur naturally. The radioisotopes typically decay into other elements upon radiating an alpha or beta particle. If an element has isotopes that are not radioactive, these are termed "stable" isotopes. All of the known stable isotopes occur naturally see primordial isotope. The many radioisotopes that are not found in nature have been characterized after being artificially made. Certain elements have no stable isotopes and are composed only of radioactive isotopes: Of the 80 elements with at least one stable isotope, 26 have only one single stable isotope. The mean number of stable isotopes for the 80 stable elements is 3. The largest number of stable isotopes that occur for a single element is 10 for tin, element Isotopic mass and atomic mass Main articles: Different isotopes of a given element are distinguished by their mass numbers, which are conventionally written as a superscript on the left hand side of the atomic symbol e. The mass number is always a whole

number and has units of "nucleons". For example, magnesium 24 is the mass number is an atom with 24 nucleons 12 protons and 12 neutrons. Whereas the mass number simply counts the total number of neutrons and protons and is thus a natural or whole number, the atomic mass of a single atom is a real number giving the mass of a particular isotope or "nuclide" of the element, expressed in atomic mass units symbol: For example, the atomic mass of chlorine to five significant digits is This number may be a fraction that is not close to a whole number. For example, the relative atomic mass of chlorine is Chemically pure and isotopically pure Chemists and nuclear scientists have different definitions of a pure element. In chemistry, a pure element means a substance whose atoms all or in practice almost all have the same atomic number , or number of protons. Nuclear scientists, however, define a pure element as one that consists of only one stable isotope. However, a pure gold ingot would be both chemically and isotopically pure, since ordinary gold consists only of one isotope, Au. Allotropy Atoms of chemically pure elements may bond to each other chemically in more than one way, allowing the pure element to exist in multiple chemical structures spatial arrangements of atoms , known as allotropes , which differ in their properties. For example, carbon can be found as diamond , which has a tetrahedral structure around each carbon atom; graphite , which has layers of carbon atoms with a hexagonal structure stacked on top of each other; graphene , which is a single layer of graphite that is very strong; fullerenes , which have nearly spherical shapes; and carbon nanotubes , which are tubes with a hexagonal structure even these may differ from each other in electrical properties. The standard state , also known as reference state, of an element is defined as its thermodynamically most stable state at a pressure of 1 bar and a given temperature typically at In thermochemistry , an element is defined to have an enthalpy of formation of zero in its standard state. For example, the reference state for carbon is graphite, because the structure of graphite is more stable than that of the other allotropes. Properties Several kinds of descriptive categorizations can be applied broadly to the elements, including consideration of their general physical and chemical properties, their states of matter under familiar conditions, their melting and boiling points, their densities, their crystal structures as solids, and their origins. General properties Several terms are commonly used to characterize the general physical and chemical properties of the chemical elements. A first distinction is between metals , which readily conduct electricity , nonmetals , which do not, and a small group, the metalloids , having intermediate properties and often behaving as semiconductors. A more refined classification is often shown in colored presentations of the periodic table. This system restricts the terms "metal" and "nonmetal" to only certain of the more broadly defined metals and nonmetals, adding additional terms for certain sets of the more broadly viewed metals and nonmetals. The version of this classification used in the periodic tables presented here includes: In this system, the alkali metals, alkaline earth metals, and transition metals, as well as the lanthanides and the actinides, are special groups of the metals viewed in a broader sense. Similarly, the reactive nonmetals and the noble gases are nonmetals viewed in the broader sense. In some presentations, the halogens are not distinguished, with astatine identified as a metalloid and the others identified as nonmetals. States of matter Another commonly used basic distinction among the elements is their state of matter phase , whether solid , liquid , or gas , at a selected standard temperature and pressure STP. Most of the elements are solids at conventional temperatures and atmospheric pressure, while several are gases. Only bromine and mercury are liquids at 0 degrees Celsius 32 degrees Fahrenheit and normal atmospheric pressure; caesium and gallium are solids at that temperature, but melt at Melting and boiling points Melting and boiling points , typically expressed in degrees Celsius at a pressure of one atmosphere, are commonly used in characterizing the various elements. While known for most elements, either or both of these measurements is still undetermined for some of the radioactive elements available in only tiny quantities. Since helium remains a liquid even at absolute zero at atmospheric pressure, it has only a boiling point, and not a melting point, in conventional presentations. Densities of the elements data page The density at a selected standard temperature and pressure STP is frequently used in characterizing the elements. Since several elements are gases at commonly encountered temperatures, their densities are usually stated for their gaseous forms; when liquefied or solidified, the gaseous elements have densities similar to those of the other elements. When an element has allotropes with different densities, one representative allotrope is typically selected in summary presentations, while densities for each allotrope can be stated where more detail is

provided. For example, the three familiar allotropes of carbon amorphous carbon , graphite , and diamond have densities of 1. Crystal structures Main article: Crystal structure The elements studied to date as solid samples have eight kinds of crystal structures: For some of the synthetically produced transuranic elements, available samples have been too small to determine crystal structures. Occurrence and origin on Earth Chemical elements may also be categorized by their origin on Earth, with the first 94 considered naturally occurring, while those with atomic numbers beyond 94 have only been produced artificially as the synthetic products of man-made nuclear reactions. Of the 94 naturally occurring elements, 83 are considered primordial and either stable or weakly radioactive. The remaining 11 naturally occurring elements possess half lives too short for them to have been present at the beginning of the Solar System , and are therefore considered transient elements. Of these 11 transient elements, 5 polonium , radon , radium , actinium , and protactinium are relatively common decay products of thorium and uranium. The remaining 6 transient elements technetium , promethium , astatine , francium , neptunium , and plutonium occur only rarely, as products of rare decay modes or nuclear reaction processes involving uranium or other heavy elements. Elements with atomic numbers 1 through 40 are all stable, while those with atomic numbers 41 through 82 except technetium and promethium are metastable. Elements with atomic numbers 83 through 94 are unstable to the point that their radioactive decay can be detected. Three of these elements, bismuth element 83 , thorium element 90 , and uranium element 92 have one or more isotopes with half-lives long enough to survive as remnants of the explosive stellar nucleosynthesis that produced the heavy elements before the formation of our solar system. For example, at over 1.

Chapter 5 : Revised Common Lectionary | Word to Worship

Like the word root, it is the basic foundation to which other word elements are added to build a complete medical word. suffix a word element placed at the end of a word that changes the meaning of the word.

Chapter 6 : element | Definition of element in English by Oxford Dictionaries

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Chapter 7 : Easy Photo & Video Editing Software | Adobe Elements Family

mark h lane calendrierdelascience.com 1 | P a g e GOD'S TRUTH IN CHEMISTRY SUMMARY God created the elements (atoms) and also assembled them into molecules.

Chapter 8 : Word for Windows training - Word

Added Word Difference Package Elements. Added Word Difference Package Elements. Added Word Field Definition. Namespace.

Chapter 9 : Francium - Element information, properties and uses | Periodic Table

In English, this seems to be what my high school English teachers would call a 'run on' sentence, but then, maybe I'm missing something crucial in understanding French word order, sentence structure, and independent/dependent clauses.